Name: Date: Period:

**Practice: Redox Reactions**

*Directions*: For each of the following reactions, determine which element is oxidized and which element is reduced.



4 Al (s) + 3 O2 (g) 🡪 2 Al2O3 (s)

Oxidized: Reduced:

2 Cr + 6 H+ 🡪 2 Cr3+ + 3 H2

Oxidized: Reduced:

2 Fe + Al2O3 🡪 Fe2O3 + 2 Al

Oxidized: Reduced:

Mg + 2 H+ 🡪 Mg2+ + H2

Oxidized: Reduced:

Br- + MnO4- 🡪 Br2 + Mn2+

Oxidized: Reduced:

Cr2O72- + Cl**-** 🡪 Cl2 + Cr3+

Oxidized: Reduced:

2 AlBr3 + 3 Cl2 🡪 2 AlCl3 + 3 Br2

Oxidized: Reduced:

Cu + S 🡪 CuS

Oxidized: Reduced:

3 H2 + N2 🡪 2 NH3

Oxidized: Reduced:

2 Na + 2 H2O 🡪 2 NaOH + H2

Oxidized: Reduced:

PbO2 + Cl- 🡪 PbCl2 + O2

Oxidized: Reduced: