Name: Date: Period:

Practice: Ideal Gas Law-Part 2

1. What pressure (in kPa) is exerted by 64.9g of Oxygen gas (O2) contained in an 4.03 L vessel at 17 ºC?
2. How many grams of Chlorine gas (Cl2) would occupy a 1.32 L flask at -19 ºC and 17.6 atm of pressure?
3. What volume is occupied by 0.579 moles of gas at 406 mmHg and 204 ºC?
4. At what pressure (in atm) is a gas if 17.96g Carbon Monoxide (CO) are contained in a 1.054 L container at 86 ºC?
5. At what temperature is a gas if 1.065 moles are in a 7.96 L vessel at 138.7 kPa? Give your answer in both K and ºC.
6. Calculate the volume of 145.7g of Sulfur Trioxide (SO3) at 68 ºC and 1.25 atm.
7. Calculate the pressure (in mmHg) of 7.2 moles Carbon Monoxide (CO) in a 9.04 L tank at 605 ºC.
8. Calculate the number of grams of a 0.107 L sample of hydrogen gas (H2) at 458 mmHg and

-75 ºC.

1. Calculate the temperature of 22.3g of nitrogen gas (N2) in a 5.09 L flask at 3.57 atm. Give your answer in both K and ºC.